

Instructions: You will only complete one step in each practice problem. After you write the equation, the next person will balance it, then the next person will write the molecular equation, and so on until all the practice problems have been solved.

Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in an aqueous solution. Include states of matter in your balanced equation.

1. Sodium chloride and lead II nitrate

Molecular Equation:

Complete Ionic Equation:

Net Ionic Equation:

Spectator ions:

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2. Lithium hydroxide and barium chloride

Molecular Equation:

Complete Ionic Equation:

Net Ionic Equation:

Spectator Ions:

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Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in an aqueous solution. Include states of matter in your balanced equation.

3. Barium Bromide and sodium sulfate

Molecular Equation:

Complete Ionic Equation:

Net Ionic Equation:

Spectator Ions:

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Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in an aqueous solution. Include states of matter in your balanced equation.

4. Potassium fluoride and magnesium nitrate

Molecular Equation:

Complete Ionic Equation:

Net Ionic Equation:

Spectator ions: