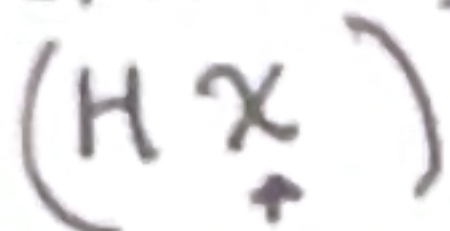


Key!

Naming Acids

Rules: Binary Acids

2 elements



some non metal

use →

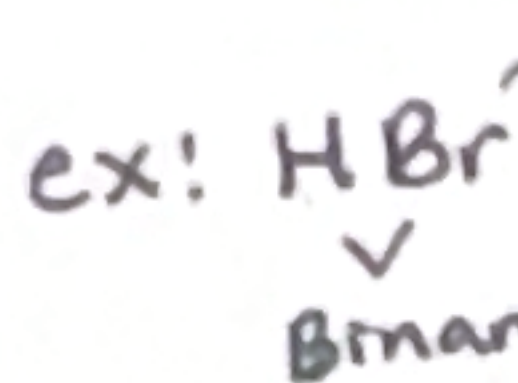
hydro

stem

-ic

suffix

acid



Bromine

hydro

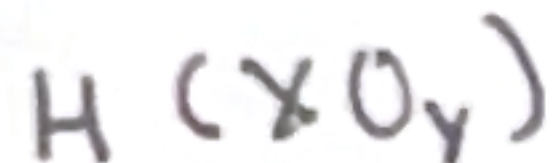
brom

ic

acid

Rules: Ternary Acids

3 elements



some

polyatomic ion

Use → if polyatomic ion ends in -ite

stem

-ous

acid

Use → if polyatomic ion ends in -ate

stem

-ic

acid

Binary →

Ternary →

Binary →

Formula	Cation	Formula for Anion	Name of Anion	Name of Acid
HCl	H ⁺	Cl ⁻	chloride	Hydrochloric acid
HNO ₃	H ⁺	NO ₃ ⁻	nitrate	Nitric acid
HF	H ⁺	F ⁻	fluoride	Hydrofluoric acid
H ₂ SO ₄	H ⁺	SO ₄ ²⁻	sulfate	Sulfuric acid stem
H ₂ CO ₃	H ⁺	CO ₃ ²⁻	carbonate	Carbonic acid
H ₂ SO ₃	H ⁺	SO ₃ ²⁻	sulfite means 1 less oxygen	Sulfurous acid
HClO ₃	H ⁺	ClO ₃ ⁻	chlorate	Chloric acid
H ₃ PO ₄	H ⁺	PO ₄ ³⁻	phosphate	Phosphoric acid
HCN	H ⁺	CN ⁻	cyanide	Hydrocyanic acid
CH ₃ COOH	H ⁺	CH ₃ COO ⁻	acetate	Acetic acid
HI	H ⁺	I ⁻	iodide	Hydroiodic acid
HClO	H ⁺	ClO ⁻	hypochlorite	Hypochlorous acid
HNO ₂	H ⁺	NO ₂ ⁻	nitrite	Nitrous acid