

Key!

Chapter 4: Reactions in Aqueous Solutions

Strong
ACIDS

HCl HNO₃
HBr H₂SO₄
HI HClO₄

Strong
Bases

LiOH Ca(OH)₂
NaOH Sr(OH)₂
KOH Ba(OH)₂

1.) Which of the following compounds is a strong electrolyte?

A.) H₂O B.) CH₃OH C.) CH₃CH₂OH D.) HF **E.) NaF**

molecular compound *molecular compound* *acid*
nonelectrolyte *nonelectrolyte* *nonelectrolyte* *weak acid*

2.) Which of the following compounds is a weak electrolyte?

A.) HNO₃ B.) NaNO₃ **C.) HNO₂** D.) NaNO₂ E.) NaOH

strong acid *strong electrolyte* *weak base*

3.) Which of the following compounds is a nonelectrolyte?

A.) NaF B.) HNO₃ C.) CH₃COOH (acetic acid) **D.) C₆H₁₂O₆ - (glucose)**

strong electrolyte *strong acid* *weak electrolyte* *molecular compound*

4.) The distinguishing characteristic of all electrolyte solutions is that they

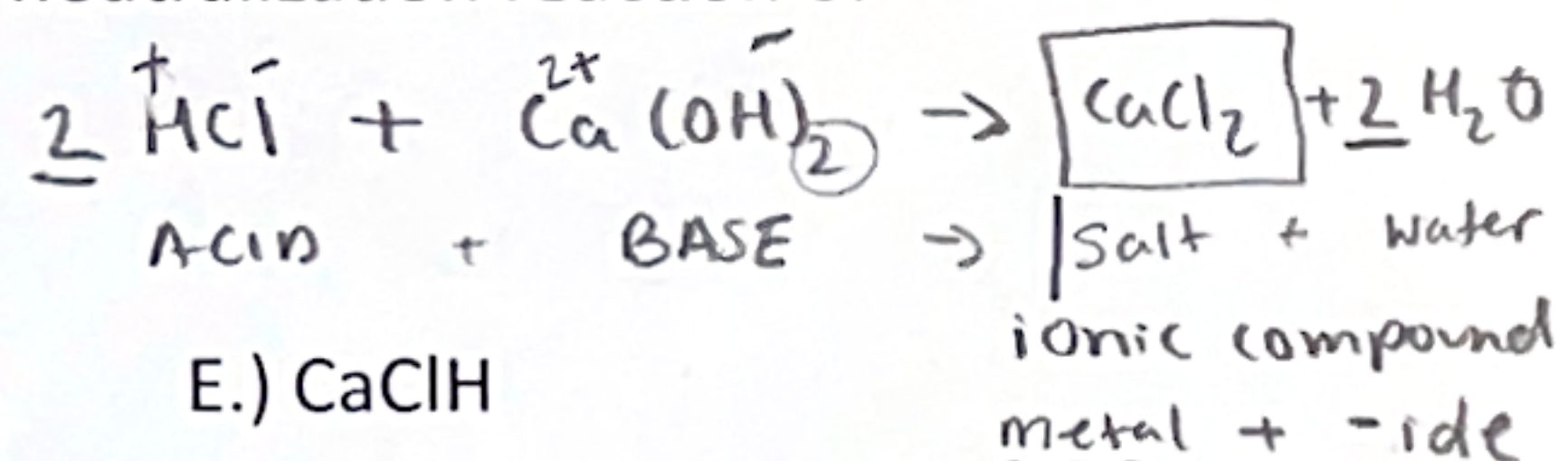
A.) Contain molecules **B.) Conduct electricity** C.) React with other solutions D.) Always contain acids E.) Conduct heat

due to process of mobile ions that can freely move when electric current is applied.

5.) What is the correct formula of the salt formed in the neutralization reaction of

hydrochloric acid with calcium hydroxide?

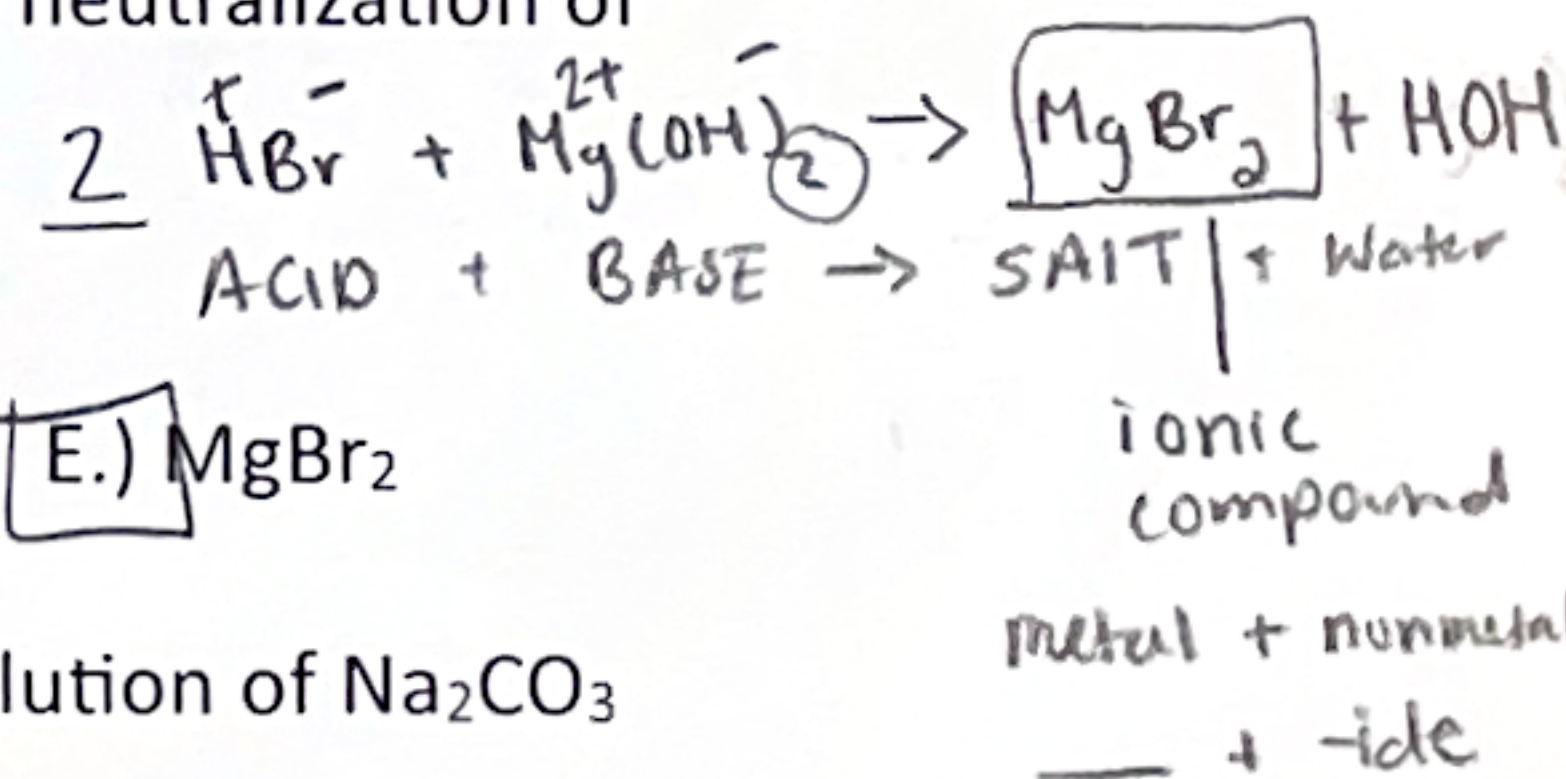
A.) CaO **B.) CaCl₂** C.) CaH₂ D.) CaCl E.) CaClH



6.) What is the chemical formula of the salt produced by the neutralization of

hydrobromic acid with magnesium hydroxide?

A.) MgBr B.) Mg₂Br₃ C.) Mg₃Br₂ D.) Mg₂Br **E.) MgBr₂**



7.) Identify the major ionic species present in an aqueous solution of Na₂CO₃

A.) Na₂⁺, CO₃²⁻ B.) Na₂⁺, C²⁻, O₃ C.) Na⁺, C⁴⁺, O₃²⁻ **D.) Na⁺, CO₃²⁻**

