

Ch. 3 Chemical Reactions & Balancing Equations

Chemical equations are used to represent chemical reactions on paper.

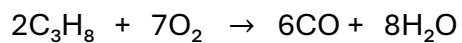
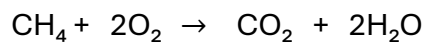
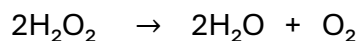
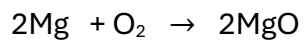
The left side of the equation is for the _____.

The right side of the equation is for the _____.

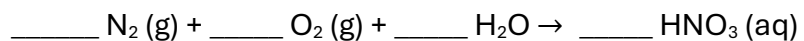
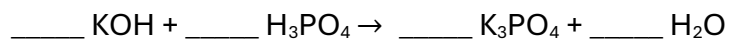
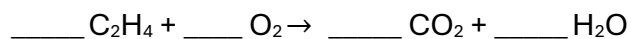
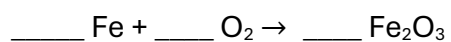
_____ represent the number of molecules or moles of that compound.

_____ indicate how many atoms of that element are in a single molecule.

Identify what type of chemical reactions are the following:



Identify the reactants and products in this reaction and balance the equation: When solid aluminum reacts with aqueous hydrochloric acid, it produces aqueous aluminum chloride and hydrogen gas.



Calculate the molecular weight of the following compounds:

Ammonia (NH_3)

Glucose ($\text{C}_6\text{H}_{12}\text{O}_6$)

Ethanol ($\text{C}_2\text{H}_5\text{OH}$)

Sulfuric Acid (H_2SO_4)

What's the formula for percent composition?

Calculate the percent composition of Calcium, Carbon, and Oxygen:

Calcium Carbonate (CaCO_3)