

Chapter 4 – Reactions

What makes a weak acid?

What are examples of strong bases?

The balanced molecular equation for the complete neutralization of HNO_3 by LiOH in aqueous solution is:

The balanced molecular equation for complete neutralization of H_2SO_4 by KOH in aqueous solution is:

When aqueous solutions of **barium nitrate ($\text{Ba}(\text{NO}_3)_2$)** and **sodium sulfate (Na_2SO_4)** are mixed, a **precipitate** forms. Write the molecular equation, complete ionic equation, and net ionic equation for this reaction.

Molecular:

Complete Ionic:

Net Ionic:

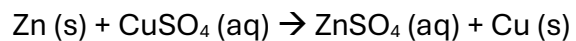
The ions formed when **MgBr₂** dissociates in water are:

The ions formed when **Na₂SO₄** dissociates in water are:

What is a redox reaction?

What happens in oxidation / reduction?

Determine which compounds are oxidized and reduced. Determine the oxidizing and reducing agents in the following reaction:



Calculate the molarity by dissolving **15.6 g of KNO₃** in **250 mL of water**.

How many grams of CaCl₂ are required to make 0.750 L of 0.500 M CaCl₂?